**11 CHEMISTRY – Term 3**

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| **TOPIC and TIMING (Weeks)** | **QCAA OBJECTIVES and ASSESSMENT** | **LEARNING GOALS and SUCCESS CRITERIA** | **Page reference** |
| **Solubility and Identifying Ions in Solution**  **Week 10 (3 lessons)** | **Unit 2:Topic 2**  **Objectives 1,2,3,4,5,6,7** | **SC97**: I can explain the relationship between the solubility of substances in water, including ionic and molecular substances, and the intermolecular forces between species in the substances and water molecules  **SC98**: I can recognise that changes in temperature can affect solubility and recall that most gases become less soluble as solvent temperature increases while most solutes become more soluble as the solvent temperature increases  **SC99**: I can interpret, analyse and evaluate data and solubility curves to communicate conceptual understanding, solve problems and make predictions  **SC100**: I can apply solubility rules to determine the products of reactions and to predict if a precipitate will form  **SC101**: I can determine the presence of specific ions in solutions based on evidence derived from chemical reactions, including precipitation and acid-carbonate reactions  **SC102**: I can construct and use appropriate representations, including ionic formulas, chemical formulas, chemical equations and phase descriptions for chemical species to communicate conceptual understanding, solve problems and make predictions  **Mandatory practical**: Precipitation reactions to identify cations and anions |  |
| **LG 1: Students can understand and explain the solubility of ionic and molecular substances in water and predict whether precipitation will occur in**  **chemical reactions.** |  |